Selected answers to homework: Lecture 11

ester
$$H_{2}C-O \xrightarrow{O} (CH_{2})_{14}CH_{3}$$

$$H_{2}C-O \xrightarrow{O} (CH_{2})_{16}CH_{3}$$

$$H_{2}C-O \xrightarrow{O} (CH_{2})_{16}CH_{3}$$

$$H_{2}C-O \xrightarrow{O} (CH_{2})_{16}CH_{3}$$

$$H_{3}CH_{2}C \xrightarrow{O} NH_{2}$$

$$H_{3$$

Propanamide and methyl acetate are both soluble in water because both are polar molecules and can hydrogen bond with water. Propanamide is higher boiling because molecules of porpanamide can hydrogen-bond with each other, as well as with water.